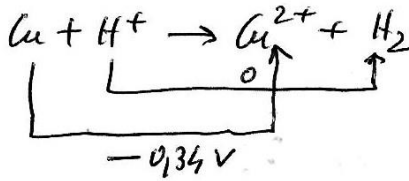


40B-PAU-REDOX.

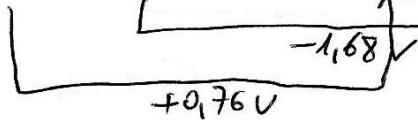
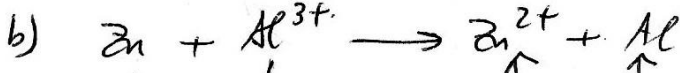
40

a)  $\text{Cu} + \text{HCl}$  <sup>ácido</sup> es lo mismo que  $\text{Cu} + \text{H}^+$



$E^\circ = -0,34 + 0 = \boxed{-0,34}$  NO REACCIONA

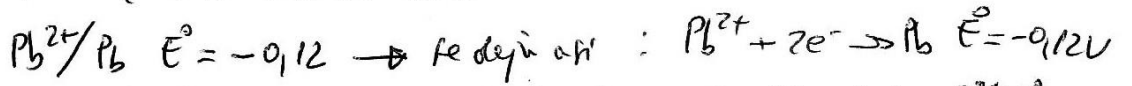
luego no se oxida



$E^\circ = -1,68 + 0,76 = \boxed{-0,92}$  NO REACCIONA

luego no se produce

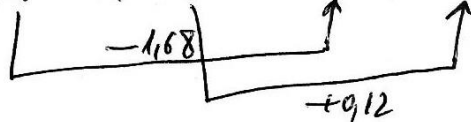
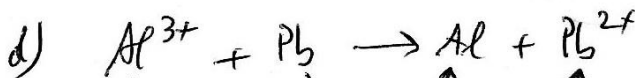
c) Pila (el  $E^\circ$  debe ser  $\oplus$ )



el Zn se oxida y el  $\text{Pb}^{2+}$  se reduce  
(ÁNODO) (CÁTODOS)

$E_{\text{pila}}^\circ = -0,12 + 0,76 = \underline{\underline{0,64}}$

Fuente:

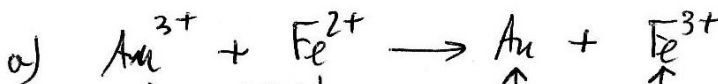


$E^\circ = -1,68 + 0,12 = \boxed{-1,56\text{V}}$

NO REACCIONA.

luego si que es ESTABLE.

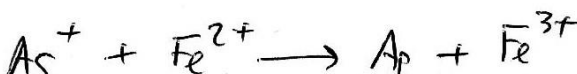
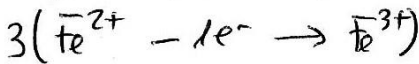
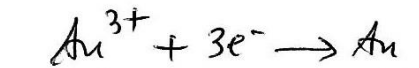
41



$E^\circ = +1,50 - 0,77 = +0,73\text{V}$

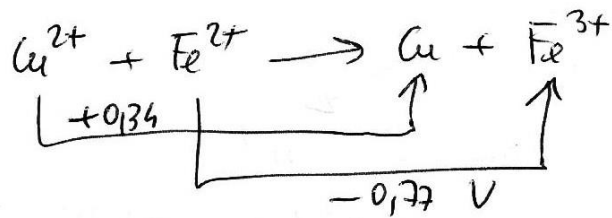
si REACCIONA.

se produce  $\text{Au} + \text{Fe}^{3+}$



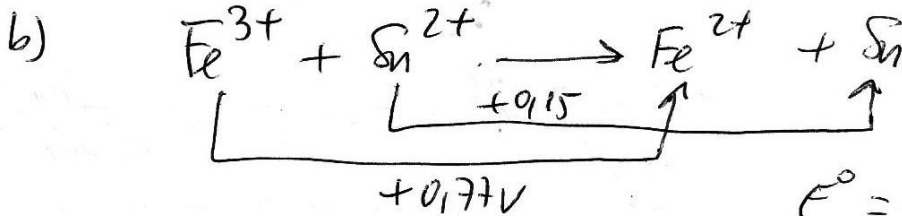
$E^\circ = +0,80 - 0,77 = \oplus$

si REACCIONA.



$$E^{\circ} = +0,34 - 0,77 = \ominus$$

NO REACCIÓ



$$E^{\circ} = +0,77 + 0,15 = \oplus$$

SI REACCIÓ

